

# Texas Class D Water License Practice Exam (Sample)

## Study Guide



**Everything you need from our exam experts!**

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**SAMPLE**

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# Introduction

Preparing for a certification exam can feel overwhelming, but with the right tools, it becomes an opportunity to build confidence, sharpen your skills, and move one step closer to your goals. At Examzify, we believe that effective exam preparation isn't just about memorization, it's about understanding the material, identifying knowledge gaps, and building the test-taking strategies that lead to success.

This guide was designed to help you do exactly that.

Whether you're preparing for a licensing exam, professional certification, or entry-level qualification, this book offers structured practice to reinforce key concepts. You'll find a wide range of multiple-choice questions, each followed by clear explanations to help you understand not just the right answer, but why it's correct.

The content in this guide is based on real-world exam objectives and aligned with the types of questions and topics commonly found on official tests. It's ideal for learners who want to:

- Practice answering questions under realistic conditions,
- Improve accuracy and speed,
- Review explanations to strengthen weak areas, and
- Approach the exam with greater confidence.

We recommend using this book not as a stand-alone study tool, but alongside other resources like flashcards, textbooks, or hands-on training. For best results, we recommend working through each question, reflecting on the explanation provided, and revisiting the topics that challenge you most.

Remember: successful test preparation isn't about getting every question right the first time, it's about learning from your mistakes and improving over time. Stay focused, trust the process, and know that every page you turn brings you closer to success.

Let's begin.

# How to Use This Guide

**This guide is designed to help you study more effectively and approach your exam with confidence. Whether you're reviewing for the first time or doing a final refresh, here's how to get the most out of your Examzify study guide:**

## **1. Start with a Diagnostic Review**

**Skim through the questions to get a sense of what you know and what you need to focus on. Your goal is to identify knowledge gaps early.**

## **2. Study in Short, Focused Sessions**

**Break your study time into manageable blocks (e.g. 30 - 45 minutes). Review a handful of questions, reflect on the explanations.**

## **3. Learn from the Explanations**

**After answering a question, always read the explanation, even if you got it right. It reinforces key points, corrects misunderstandings, and teaches subtle distinctions between similar answers.**

## **4. Track Your Progress**

**Use bookmarks or notes (if reading digitally) to mark difficult questions. Revisit these regularly and track improvements over time.**

## **5. Simulate the Real Exam**

**Once you're comfortable, try taking a full set of questions without pausing. Set a timer and simulate test-day conditions to build confidence and time management skills.**

## **6. Repeat and Review**

**Don't just study once, repetition builds retention. Re-attempt questions after a few days and revisit explanations to reinforce learning. Pair this guide with other Examzify tools like flashcards, and digital practice tests to strengthen your preparation across formats.**

**There's no single right way to study, but consistent, thoughtful effort always wins. Use this guide flexibly, adapt the tips above to fit your pace and learning style. You've got this!**

## Questions

- 1. What can be a consequence of closing a hydrant valve too quickly?**
  - A. Leakage**
  - B. Water hammer**
  - C. Air lock**
  - D. Backflow**
- 2. The minimum number of samples required is largely influenced by which aspect of the water system?**
  - A. The length of the distribution lines**
  - B. The population served**
  - C. The source of water**
  - D. The type of treatment employed**
- 3. Water is classified as hard when it exceeds how many mg/L of calcium carbonate?**
  - A. 50 mg/L**
  - B. 100 mg/L**
  - C. 150 mg/L**
  - D. 200 mg/L**
- 4. How many years must bacteriological sample results be retained?**
  - A. 3 years**
  - B. 5 years**
  - C. 7 years**
  - D. 10 years**
- 5. What is the function of a chemical pump in a hypochlorinator system?**
  - A. To store chemicals**
  - B. To mix water with chemicals**
  - C. To dispense the chemical solution**
  - D. To measure the chemical concentration**

- 6. Sodium hypochlorite is commonly referred to as what?**
- A. Liquid chlorine**
  - B. Liquid bleach**
  - C. Powdered bleach**
  - D. Calcium chloride**
- 7. What is the MCL for total trihalomethanes in water systems using surface water?**
- A. 0.06 mg/L**
  - B. 0.08 mg/L**
  - C. 0.10 mg/L**
  - D. 0.04 mg/L**
- 8. What must be instituted when normal chlorine levels are not maintained in the water system?**
- A. No action is required**
  - B. Emergency treatment protocols**
  - C. Safety inspections**
  - D. Public education programs**
- 9. What is a primary function of storage tanks in water treatment systems?**
- A. Uniformity of pump rates**
  - B. Supply water for firefighting**
  - C. Provide time for disinfection**
  - D. All of the above**
- 10. What needs to be controlled in order to maintain well integrity and water quality?**
- A. Water temperature**
  - B. Water levels**
  - C. Contamination sources**
  - D. Well pump efficiency**



## **Answers**

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1. B
2. B
3. B
4. B
5. C
6. B
7. B
8. B
9. D
10. C

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## **Explanations**

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**1. What can be a consequence of closing a hydrant valve too quickly?**

**A. Leakage**

**B. Water hammer**

**C. Air lock**

**D. Backflow**

Closing a hydrant valve too quickly can lead to the phenomenon known as water hammer. Water hammer occurs when there is a sudden change in the flow of water, which causes a pressure surge in the piping system. This pressure surge can result in loud banging noises and potentially damaging vibrations within the pipes. When a hydrant valve is closed rapidly, the momentum of the flowing water does not have time to adjust to the sudden stop, creating a hydraulic shock wave that propagates back through the system. This can lead to stress on pipes, joints, and fittings, and in severe cases, may cause ruptures or leaks. While leakage, air locks, and backflow are certainly important considerations in water system management, they are not direct consequences of rapidly closing a hydrant valve. Leakage may occur from other issues such as wear and tear or improper installation, while air locks typically arise from trapped air in the system. Backflow is related to the reverse flow of water due to differences in pressure but is not an effect of closing a valve too quickly. Therefore, water hammer clearly stands out as the immediate and relevant consequence in this context.

**2. The minimum number of samples required is largely influenced by which aspect of the water system?**

**A. The length of the distribution lines**

**B. The population served**

**C. The source of water**

**D. The type of treatment employed**

The minimum number of samples required is primarily influenced by the population served by the water system. This relationship is established in water quality standards and regulations, which mandate specific sampling frequencies and volumes based on the number of people receiving water from the system. The rationale behind this is that a larger population presents a greater potential for contamination and variability in water quality, necessitating more rigorous monitoring. Sampling is essential to ensure that the water supplied meets safety and health standards. For example, a community supplying water to hundreds of people must conduct more frequent and numerous tests than a system serving a small number of individuals. This scaling of sampling requirements helps ensure that any potential issues with water quality are identified and addressed promptly, thereby protecting public health. While factors like the length of distribution lines, the source of water, or the type of treatment employed can affect water quality, they do not have as direct an impact on the sample size requirements set forth by regulatory guidelines compared to the population served.

**3. Water is classified as hard when it exceeds how many mg/L of calcium carbonate?**

**A. 50 mg/L**

**B. 100 mg/L**

**C. 150 mg/L**

**D. 200 mg/L**

Water is classified as hard when it contains more than 100 mg/L of calcium carbonate. This classification is crucial in water treatment and quality assessment, as hard water can cause scaling in pipes and water heaters, making it less suitable for certain uses. Specifically, hardness in water is primarily due to the presence of calcium and magnesium ions, and their concentration is measured as equivalent calcium carbonate. In terms of managing water resources, understanding the hardness of water informs treatment processes needed for domestic and industrial uses, as well as impacts on soap and detergent effectiveness. Water that exceeds 100 mg/L may be significantly impacted by these concerns, making this threshold an important benchmark in water quality studies.

**4. How many years must bacteriological sample results be retained?**

**A. 3 years**

**B. 5 years**

**C. 7 years**

**D. 10 years**

Bacteriological sample results must be retained for a minimum of 5 years. This retention period is established to ensure that there is a historical record of water quality data, which is vital for monitoring trends, ensuring compliance with public health regulations, and addressing any potential issues that may arise regarding water safety and quality. Retaining these records for five years allows for sufficient time to analyze patterns in water quality, assist in regulatory reporting, and facilitate proper management of water supply systems. In the context of regulatory requirements, a five-year retention period also aligns with the guidelines set forth by various environmental protection agencies, ensuring that water utilities can provide necessary documentation in case of audits or health investigations. Keeping these records readily accessible helps promote transparency and accountability within water management practices, crucial for public health trust in water supplies.

**5. What is the function of a chemical pump in a hypochlorinator system?**

- A. To store chemicals**
- B. To mix water with chemicals**
- C. To dispense the chemical solution**
- D. To measure the chemical concentration**

In a hypochlorinator system, the primary function of a chemical pump is to dispense the chemical solution, specifically chlorine or other disinfectants, into the water system. This ensures that the correct amount of chlorine is delivered at the desired concentration for effective disinfection. Proper dosing is crucial for controlling pathogens in water supplies, maintaining water quality, and ensuring public health safety. While other components may handle storing chemicals or mixing, those functions do not involve direct application into the water. A chemical pump's role is focused on accurately and consistently releasing the chemical solution into the flow of water, which facilitates timely and effective disinfection processes. Measurements of chemical concentration or mixing may be parts of a larger system, but they are not the primary role of the pump itself.

**6. Sodium hypochlorite is commonly referred to as what?**

- A. Liquid chlorine**
- B. Liquid bleach**
- C. Powdered bleach**
- D. Calcium chloride**

Sodium hypochlorite is commonly known as liquid bleach due to its widespread use as a disinfectant and a bleaching agent. In water treatment and sanitation processes, sodium hypochlorite is often used to eliminate pathogens, purify water, and remove unwanted colors or stains from surfaces. The liquid form makes it easy to handle and dilute for various applications, which is why it is recognized as a form of bleach. Liquid chlorine, on the other hand, refers specifically to chlorine gas that has been liquefied for storage and transport, while powdered bleach usually pertains to calcium hypochlorite, which is a solid form of chlorine compound used for disinfection. Calcium chloride is an entirely different chemical compound used primarily for de-icing and as a drying agent, making it unrelated to the context of this question.

**7. What is the MCL for total trihalomethanes in water systems using surface water?**

- A. 0.06 mg/L
- B. 0.08 mg/L**
- C. 0.10 mg/L
- D. 0.04 mg/L

The maximum contaminant level (MCL) for total trihalomethanes (TTHMs) in water systems using surface water is set at 0.08 mg/L. This standard is established to protect public health, as TTHMs are known byproducts of chlorination used in water treatment, and their presence in drinking water can have adverse health effects over time, including potential risks of cancer. The regulation aims to ensure that the levels of these compounds are kept within safe limits, thereby safeguarding the health of consumers relying on treated surface water supplies. Understanding this standard is essential for water system operators to maintain compliance and ensure the safety of the water supply.

**8. What must be instituted when normal chlorine levels are not maintained in the water system?**

- A. No action is required
- B. Emergency treatment protocols**
- C. Safety inspections
- D. Public education programs

When normal chlorine levels are not maintained in the water system, it is essential to institute emergency treatment protocols. Chlorine is a critical component in water treatment as it acts as a disinfectant to eliminate harmful microorganisms that can pose serious health risks to consumers. If chlorine levels drop below acceptable levels, it indicates a potential failure in the disinfection process, which can lead to contamination of the water supply. Emergency treatment protocols involve immediate actions to restore proper chlorination and ensure the safety of the water. These might include increasing chlorine dosage, conducting additional tests to identify contamination, and possibly issuing boil water advisories or providing alternative safe water supplies until the issue is resolved. Such protocols are vital to protect public health and prevent disease outbreaks related to waterborne pathogens. In contrast, simply taking no action or relying on safety inspections or public education programs would not address the immediate risk posed by inadequate chlorination. When water safety is compromised, rapid response measures are crucial for maintaining the integrity of the water supply system and safeguarding the health of the community.

**9. What is a primary function of storage tanks in water treatment systems?**

- A. Uniformity of pump rates**
- B. Supply water for firefighting**
- C. Provide time for disinfection**
- D. All of the above**

In water treatment systems, storage tanks serve several critical functions that contribute to the overall efficiency and effectiveness of the water supply process. One important role is to maintain uniformity of pump rates, which helps ensure a steady and consistent flow of water through the system. This uniformity is crucial for maintaining pressure and preventing surges or drops that can affect both treatment processes and distribution. Additionally, storage tanks play a vital role in providing water for firefighting. By having a designated supply of water available in storage tanks, municipalities can ensure that there is adequate water pressure and volume readily accessible for emergency response scenarios. Another key function of storage tanks is to provide time for disinfection processes to occur. After the water undergoes treatment, holding it in storage tanks allows for the necessary contact time to achieve effective disinfection. This is essential for removing pathogens and ensuring that the water is safe for consumption. Given these multiple functions—maintaining pump rates, supplying water for emergencies, and allowing for proper disinfection—it's clear why the answer encompasses all of these aspects. Each function contributes to the safety, reliability, and efficiency of the water treatment operation, making the choice that includes all of these elements the most comprehensive answer.

**10. What needs to be controlled in order to maintain well integrity and water quality?**

- A. Water temperature**
- B. Water levels**
- C. Contamination sources**
- D. Well pump efficiency**

Maintaining well integrity and water quality is primarily focused on preventing contaminants from entering the water supply. This requires controlling contamination sources. When there are unregulated or unmanaged sources of contamination near a well, pollutants can seep into the aquifer, thereby degrading the quality of the water drawn from the well. Effective management practices involve identifying potential sources of contamination, such as agricultural runoff, sewage discharge, improper waste disposal, and industrial activities. By addressing these sources, water systems can enhance the safety and purity of the water supply, ensuring the health of consumers and the environment. While other factors like water temperature, water levels, and well pump efficiency can impact the operation of a well, they do not directly relate to the integrity of the well structure or the safeguarding of the water quality against contamination. Therefore, focusing on contamination sources is essential in the context of overall water quality management.



## Next Steps

**Congratulations on reaching the final section of this guide. You've taken a meaningful step toward passing your certification exam and advancing your career.**

**As you continue preparing, remember that consistent practice, review, and self-reflection are key to success. Make time to revisit difficult topics, simulate exam conditions, and track your progress along the way.**

**If you need help, have suggestions, or want to share feedback, we'd love to hear from you. Reach out to our team at [hello@examzify.com](mailto:hello@examzify.com).**

**Or visit your dedicated course page for more study tools and resources:**

**<https://texasclassdwater.examzify.com>**

**We wish you the very best on your exam journey. You've got this!**