

HESI A2 Chemistry Practice Test (Sample)

Study Guide



Everything you need from our exam experts!

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Introduction

Preparing for a certification exam can feel overwhelming, but with the right tools, it becomes an opportunity to build confidence, sharpen your skills, and move one step closer to your goals. At Examzify, we believe that effective exam preparation isn't just about memorization, it's about understanding the material, identifying knowledge gaps, and building the test-taking strategies that lead to success.

This guide was designed to help you do exactly that.

Whether you're preparing for a licensing exam, professional certification, or entry-level qualification, this book offers structured practice to reinforce key concepts. You'll find a wide range of multiple-choice questions, each followed by clear explanations to help you understand not just the right answer, but why it's correct.

The content in this guide is based on real-world exam objectives and aligned with the types of questions and topics commonly found on official tests. It's ideal for learners who want to:

- Practice answering questions under realistic conditions,
- Improve accuracy and speed,
- Review explanations to strengthen weak areas, and
- Approach the exam with greater confidence.

We recommend using this book not as a stand-alone study tool, but alongside other resources like flashcards, textbooks, or hands-on training. For best results, we recommend working through each question, reflecting on the explanation provided, and revisiting the topics that challenge you most.

Remember: successful test preparation isn't about getting every question right the first time, it's about learning from your mistakes and improving over time. Stay focused, trust the process, and know that every page you turn brings you closer to success.

Let's begin.

How to Use This Guide

This guide is designed to help you study more effectively and approach your exam with confidence. Whether you're reviewing for the first time or doing a final refresh, here's how to get the most out of your Examzify study guide:

1. Start with a Diagnostic Review

Skim through the questions to get a sense of what you know and what you need to focus on. Your goal is to identify knowledge gaps early.

2. Study in Short, Focused Sessions

Break your study time into manageable blocks (e.g. 30 - 45 minutes). Review a handful of questions, reflect on the explanations.

3. Learn from the Explanations

After answering a question, always read the explanation, even if you got it right. It reinforces key points, corrects misunderstandings, and teaches subtle distinctions between similar answers.

4. Track Your Progress

Use bookmarks or notes (if reading digitally) to mark difficult questions. Revisit these regularly and track improvements over time.

5. Simulate the Real Exam

Once you're comfortable, try taking a full set of questions without pausing. Set a timer and simulate test-day conditions to build confidence and time management skills.

6. Repeat and Review

Don't just study once, repetition builds retention. Re-attempt questions after a few days and revisit explanations to reinforce learning. Pair this guide with other Examzify tools like flashcards, and digital practice tests to strengthen your preparation across formats.

There's no single right way to study, but consistent, thoughtful effort always wins. Use this guide flexibly, adapt the tips above to fit your pace and learning style. You've got this!

Questions

1. What is the chemical formula of hydrochloric acid?
 - A. H_2O
 - B. NaCl
 - C. HCl
 - D. CH_3COOH
2. Which type of bond involves the attraction between oppositely charged ions?
 - A. Ionic bond
 - B. Covalent bond
 - C. Metallic bond
 - D. All of the above
3. What is the main component of vinegar?
 - A. Citric acid
 - B. Acetic acid
 - C. Formic acid
 - D. Lactic acid
4. What is the chemical formula for ammonia?
 - A. NH_4
 - B. NH_2
 - C. NH_3
 - D. N_2H_6
5. Which type of chemical reaction releases energy?
 - A. Endothermic reaction
 - B. Exothermic reaction
 - C. Combustion reaction
 - D. Decomposition reaction
6. What is defined as the average mass of the isotopes of an element?
 - A. Atomic Weight
 - B. Molar Mass
 - C. Atomic Number
 - D. Molecular Weight

- 7. What forces are indicated by temporary dipoles created by the movement of electrons?**
- A. Dipole-dipole interactions**
 - B. Dispersion forces**
 - C. Covalent forces**
 - D. Electrostatic forces**
- 8. Which group of elements is known as noble gases?**
- A. Group 1A**
 - B. Group 2A**
 - C. Group 7A**
 - D. Group 8A**
- 9. Which of the following best describes ionic bonds?**
- A. Formed through sharing of electrons**
 - B. Formed through loss and gain of electrons**
 - C. Weak electrostatic attraction between neutral molecules**
 - D. Always involve nonmetals**
- 10. Describe the process of "chromatography".**
- A. A process of heating substances for separation**
 - B. A technique for separating components of a mixture based on their affinity for a stationary phase**
 - C. A method for determining the concentration of a solution**
 - D. A chemical reaction that produces new compounds**

Answers

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1. C
2. A
3. B
4. C
5. B
6. A
7. B
8. D
9. B
10. B

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Explanations

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1. What is the chemical formula of hydrochloric acid?

- A. H₂O
- B. NaCl
- C. HCl**
- D. CH₃COOH

Hydrochloric acid is a strong acid composed of hydrogen and chlorine. The chemical formula for hydrochloric acid is represented as HCl, which denotes that each molecule of hydrochloric acid consists of one hydrogen (H) atom bonded to one chlorine (Cl) atom. This simple structure makes it a diatomic molecule, and its acidity is due to the ability of the hydrogen ion (H⁺) to dissociate and contribute to the acidic properties in aqueous solutions. The other compounds listed in the choices serve different purposes: H₂O is the chemical formula for water, a vital substance for all known forms of life; NaCl is the formula for sodium chloride, commonly known as table salt; and CH₃COOH is acetic acid, a weaker organic acid found in vinegar. Therefore, the correct choice accurately identifies hydrochloric acid, which is essential in various chemical reactions and processes.

2. Which type of bond involves the attraction between oppositely charged ions?

- A. Ionic bond**
- B. Covalent bond
- C. Metallic bond
- D. All of the above

The attraction between oppositely charged ions defines an ionic bond. This type of bond forms when an electron is transferred from one atom to another, resulting in the formation of ions. One atom becomes positively charged (a cation) while the other becomes negatively charged (an anion). The electrostatic force between these oppositely charged ions creates a strong bond, which is characteristic of ionic compounds. In contrast, a covalent bond involves the sharing of electrons between atoms rather than the transfer of electrons, leading to neutral atoms rather than charged ions. Metallic bonds, on the other hand, occur between metal atoms, where electrons are shared in a "sea of electrons," allowing for properties like conductivity and malleability but not involving ionic attraction. Therefore, the definition and formation process of an ionic bond clearly align with the description of attraction between oppositely charged ions, making that the correct choice.

3. What is the main component of vinegar?

- A. Citric acid
- B. Acetic acid**
- C. Formic acid
- D. Lactic acid

The main component of vinegar is acetic acid, which is responsible for its characteristic sour taste and strong smell. Acetic acid is formed during the fermentation process when ethanol, a type of alcohol, is oxidized by acetic acid bacteria. This fermentation can occur naturally or may be initiated intentionally in the production of vinegar. The concentration of acetic acid in commercial vinegar typically ranges from 4% to 8% by volume. Vinegar is widely used in cooking, food preservation, and as a condiment due to its unique flavor profile. Understanding that acetic acid is the key component helps highlight why vinegar is effective in culinary applications, as its acidity can enhance flavors and also acts as an antimicrobial agent, contributing to food preservation. The other acids listed—citric acid, formic acid, and lactic acid—are indeed acidic compounds but are not responsible for the primary characteristics of vinegar. Citric acid is commonly found in citrus fruits, formic acid is associated with ants and some plants, and lactic acid is produced during the fermentation of sugars in dairy products. Each has unique properties and uses but does not define vinegar's composition.

4. What is the chemical formula for ammonia?

- A. NH_4
- B. NH_2
- C. NH_3**
- D. N_2H_6

Ammonia is a compound consisting of nitrogen and hydrogen, with the chemical formula representing one nitrogen atom bonded to three hydrogen atoms. This is distinctly reflected in the formula NH_3 , where the "N" stands for nitrogen and the three "H" symbols indicate the three hydrogen atoms that are covalently bonded to the nitrogen atom. In the context of the other options, NH_4 represents ammonium, which is a positively charged ion (cation) derived from ammonia but has a different structure. NH_2 describes the amide functional group, which does not encompass the full tetrahedral geometry or the characteristic properties of ammonia. N_2H_6 , on the other hand, suggests a molecular structure containing two nitrogen atoms and six hydrogen atoms, which is not representative of ammonia but rather a different compound called hydrazine. Thus, the formula NH_3 correctly identifies ammonia and reflects its molecular composition accurately.

5. Which type of chemical reaction releases energy?

- A. Endothermic reaction
- B. Exothermic reaction**
- C. Combustion reaction
- D. Decomposition reaction

A chemical reaction that releases energy is identified as an exothermic reaction. During this type of reaction, the total energy of the products is less than that of the reactants. As the reaction takes place, energy is released, often in the form of heat or light. This energy release occurs because the formation of new bonds in the products releases more energy than is required to break the bonds in the reactants. Common examples of exothermic reactions include combustion and certain types of oxidation reactions. Although combustion reactions, which are a specific type of exothermic reaction, also release energy, not all exothermic reactions are combustion reactions. Similarly, while decomposition reactions can be endothermic (absorbing energy), this does not apply to all decomposition reactions, making it essential to distinguish between different types of reactions based on their energy profiles. Thus, the classification of an exothermic reaction encompasses a wider range of reactions where energy is released.

6. What is defined as the average mass of the isotopes of an element?

- A. Atomic Weight**
- B. Molar Mass
- C. Atomic Number
- D. Molecular Weight

The average mass of the isotopes of an element refers to atomic weight, which takes into consideration the relative abundances of each isotope and their respective masses. Each isotope of an element has a different mass owing to variations in the number of neutrons in the atomic nucleus. The atomic weight is calculated by multiplying the mass of each isotope by its natural abundance (as a fraction of 1), and then summing these values. This value is often expressed in atomic mass units (amu) and provides a weighted average that reflects the isotopic composition of the element as it occurs in nature. In contrast, molar mass refers to the mass of one mole of a substance and is numerically equivalent to the atomic weight but is expressed in grams per mole. Atomic number refers to the number of protons in the nucleus of an atom, which defines the element itself but does not give information about isotopes. Molecular weight pertains to the sum of the atomic weights of the atoms in a molecule and is relevant for compounds rather than individual elements.

7. What forces are indicated by temporary dipoles created by the movement of electrons?

A. Dipole-dipole interactions

B. Dispersion forces

C. Covalent forces

D. Electrostatic forces

The correct answer is dispersion forces, which are a type of intermolecular force that arise due to the temporary dipoles created by the movement of electrons in atoms or molecules. As electrons are in constant motion, at any given moment, they can accumulate unevenly around the nucleus. This creates a temporary dipole, where one side of the atom becomes slightly more negative, and the other side becomes slightly more positive. These temporary dipoles can induce similar dipoles in neighboring atoms or molecules, leading to an attraction between them. This attraction is what is referred to as dispersion forces, also known as London dispersion forces, and it is a fundamental interaction in all types of matter, significantly influencing the properties of nonpolar substances. Other types of intermolecular forces, such as dipole-dipole interactions and covalent forces, involve different mechanisms and structures. For instance, dipole-dipole interactions occur between polar molecules that have permanent dipoles, while covalent forces pertain to the sharing of electrons between atoms in a molecule. Electrostatic forces generally refer to the forces between charged particles or ions and are not applicable in this context where we are specifically discussing the temporary dipoles due to electron movement.

8. Which group of elements is known as noble gases?

A. Group 1A

B. Group 2A

C. Group 7A

D. Group 8A

Noble gases are a group of elements found in Group 8A of the periodic table. This group includes helium, neon, argon, krypton, xenon, and radon. The defining characteristic of noble gases is that they have a full valence shell, which makes them highly unreactive under normal conditions. This stability is attributed to having the maximum number of electrons in their outermost energy level, resulting in a lack of tendency to gain or lose electrons. The other groups listed do not have the same properties. For instance, Group 1A consists of alkali metals, which are very reactive and have only one electron in their outer shell. Group 2A contains alkaline earth metals that are also reactive, but they have two electrons in their outer shell. Lastly, Group 7A contains halogens, which are highly reactive nonmetals with seven valence electrons, and they tend to gain one electron to achieve a stable configuration. Thus, the noble gases' unique stability and inertness distinctly set them apart from the other groups.

9. Which of the following best describes ionic bonds?

- A. Formed through sharing of electrons
- B. Formed through loss and gain of electrons**
- C. Weak electrostatic attraction between neutral molecules
- D. Always involve nonmetals

Ionic bonds are best described as being formed through the loss and gain of electrons. In ionic bonding, one atom, typically a metal, donates one or more of its electrons to another atom, usually a nonmetal. This transfer of electrons allows the metal to achieve a stable electron configuration, often resembling that of the nearest noble gas. The atom that loses electrons becomes positively charged, forming a cation, while the atom that gains electrons becomes negatively charged, forming an anion. The electrostatic attraction between these oppositely charged ions creates the ionic bond. This process differs significantly from covalent bonding, where electrons are shared between atoms, hence the first option does not accurately describe ionic bonds. The third option refers to weak attractions, which describe van der Waals forces or hydrogen bonding rather than ionic bonds. The final statement about ionic bonds always involving nonmetals is misleading; while nonmetals are often part of ionic compounds, the bond itself is between a metal and a nonmetal, not exclusively involving nonmetals.

10. Describe the process of "chromatography".

- A. A process of heating substances for separation
- B. A technique for separating components of a mixture based on their affinity for a stationary phase**
- C. A method for determining the concentration of a solution
- D. A chemical reaction that produces new compounds

Chromatography is a technique used to separate components of a mixture based on their differing affinities for a stationary phase and a mobile phase. In this process, a mixture is passed through a medium (the stationary phase), such as a column packed with silica or a thin layer of a material, while a solvent (the mobile phase) carries the mixture through this medium. The components of the mixture interact with the stationary phase to varying degrees. As a result, some components move more slowly due to stronger interactions, while others travel faster through the stationary phase. This differential movement leads to the separation of the components within the mixture, which can then be collected and analyzed. Other methods mentioned, such as heating substances for separation, determining concentration, or conducting chemical reactions, do not accurately describe the fundamental mechanism of chromatography. The essence of chromatography lies in the separation based on the unique characteristics of each component's interaction with the stationary phase, making option B the correct portrayal of the chromatography process.

Next Steps

Congratulations on reaching the final section of this guide. You've taken a meaningful step toward passing your certification exam and advancing your career.

As you continue preparing, remember that consistent practice, review, and self-reflection are key to success. Make time to revisit difficult topics, simulate exam conditions, and track your progress along the way.

If you need help, have suggestions, or want to share feedback, we'd love to hear from you. Reach out to our team at hello@examzify.com.

Or visit your dedicated course page for more study tools and resources:

<https://hesia2chemistry.examzify.com>

We wish you the very best on your exam journey. You've got this!